The low solubility of **7 in** the usual polar electrochemistry solvents (acetonitrile, methylene chloride, etc.) has, so far, thwarted our attempts to grow salts electrochemically. We have, however, been able to prepare a TCNQ salt but so far only in the form of microcrystalline powder. Preliminary ESR measurements on this solid give a g value of 2.0039 and a line width of 4 G, indicating that the spins seem to be observable only on the TCNQ molecules at 4 K. To date we measured only compressed powders of the TCNQ salt and found a resistivity of 0.6-1 Ω cm, comparable to TTF TCNQ.

We are currently attempting to improve crystal growth of the TCNQ salt as well as prepare other salts of HMTTeF and synthesize TMTTeF.[26]

Acknowledgment. We thank S. H. Bertz for helpful discussions and some initial work related to cyclopentene dianion, E. A. Chandross for helpful suggestions and critical reading of the manuscript, and R. S. Hutton and W. M. Walsh for solid-state ESR measurements.

Registry No. 2, 75415-78-0; **3**, 23586-46-1; **4**, 80631-57-8; **5**, 80631-58-9; **6**, 80631-59-0; **7**, 80631-60-3; **7** TCNQ salt, 80631-61-4.

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(26) A nonfulvenoid organic compound containing four tellurium atoms has recently been reported: D. J. Sandman, J. C. Stark, G. P. Hamill, W. A. Burke, B. M. Foxman, International Conference on Low Dimensional Conductors, Boulder, CO, August 9-14, 1981. See also: Shibaeva, R. P.; Kaminskii, V. F. *Cryst. Struct. Commun.* **1981**, *10*, 663.

(27) This weakening of the endocyclic C=C bond is probably due to transannular nonbonded Te-Te interactions in the five-membered ring and across rings. A six-membered ring isomer[28] would be expected to be less strained and consequently not show as dramatic a shift.

(28) A referee has correctly pointed out a fact, which has also not escaped our notice, that the synthetic method could also yield the six-membered ring isomer. The only way to resolve this structural problem is via X-ray crystallography. Dr. E. M. Engler has kindly supplied us with a sample of HMTSF·TCNQ for this purpose.

Additions and Corrections

Total Syntheses of (±)-Daunomycinone: Regiospecific Preparations of (±)-7,9-Dideoxydaunomycinone and 6,11-Dihydroxy-4-methoxy-7,8,9,10-tetrahydronaphthacene-5,9,12-trione [*J. Am. Chem. Soc.* **1981**, *103*, 6378]. FRANK M. HAUSER* and SUBBARAO PRASANNA, Department of Chemistry and Biological Sciences, The Oregon Graduate Center, Beaverton, Oregon 97006.

Page 6380: The three lines under Scheme II beginning with endo-18, exo-19,20, and endo-4 do not belong with this paper.

Carbon-14 Kinetic Isotope Effects in the Menschutkin-Type Reaction of Benzyl Benzenesulfonates with *N*,*N*-Dimethylanilines. Variation of the Effects with Substituents [*J. Am. Chem. Soc.* **1979**, *101*, 266]. HIROSHI YAMATAKA and TAKASHI ANDO,* The Institute of Scientific and Industrial Research, Osaka University, Suita, Osaka 565, Japan.

The reaction temperature shown in eq 1 and Table I should be 35 °C.

Catalysis of Ester Aminolysis by Divalent Metal Ions [*J. Am. Chem. Soc.* **1981**, *103*, 3233-3235]. R. M. PROPST III and L. S. TRZUPEK,* Department of Chemistry, Furman University, Greenville, South Carolina 29613.

Page 3234: The correct value for the second-order rate coefficient, k, in eq 1 is 1.5×10^{-4} M^{-1} s^{-1}, not 1.5×10^{-1} as shown.

Facile Reduction of Carbon Dioxide by Anionic Group 6B Metal Hydrides. Chemistry Relevant to Catalysis of the Water-Gas Shift Reaction [*J. Am. Chem. Soc.* **1981**, *103*, 3223]. DONALD J. DARENSBOURG,* ANDRZEJ ROKICKI, and MARCETTA Y. DARENSBOURG, Department of Chemistry, Tulane University, New Orleans, Louisiana 70118.

Page 3223, Table I: Two sets of numbers were interchanged. The ^{13}C NMR data for $HC(O)OCr(CO)_5^-$ is $\delta C(eq) = 216.0$ and $\delta C(ax) = 222.7$; for $HC(S)SCr(CO)_5^-$ the $\delta C(eq) = 218.0$ and the $\delta C(ax) = 225.4$ ppm.

Optically Detected Magnetic Resonance Evidence for Carcinogen-Nucleic Acid Interaction in the Tetrahydro-9,10-epoxybenzo[*e*]pyrene-DNA Adduct [*J. Am. Chem. Soc.* **1981**, *103*, 5257-5259]. STEVEN M. LEFKOWITZ and HENRY C. BRENNER,* Department of Chemistry, New York University, New York, New York 10003.

Page 5258, left column, first paragraph: The limits of error in the ODMR frequencies should be ±2 MHz (megahertz), rather than ±2 GHz, as printed. The ODMR frequencies themselves are in GHz.

Transition-Metal Complexes of Vitamin B_6 Related Compounds. 3. X-ray, Mössbauer, and Magnetic Properties of a Binuclear Iron(III) Complex Containing an Unusual Pyridoxal Derivative [*J. Am. Chem. Soc.* **1980**, *102*, 6040]. GARY J. LONG,* JAMES T. WROBLESKI, RAJU V. THUNDATHIL, DON M. SPARLIN, and E. O. SCHLEMPER, Departments of Chemistry, University of Missouri—Rolla, Rolla, Missouri 65401, and University of Missouri—Columbia, Columbia, Missouri 65201.

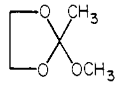
The space group and the extinctions reported in the abstract and on page 6041 are in error. The correct space group is *Pbca* and the correct extinctions are odd k absent for $0kl$, and l absent for $h0l$, and odd h absent for $hk0$. The structure was refined in the correct space group and all the structural results presented in the paper are correct.

Structure of Palytoxin [*J. Am. Chem. Soc.* **1981**, *103*, 2491-2494]. RICHARD E. MOORE* and GIOVANNI BARTOLINI, Department of Chemistry, University of Hawaii, Honolulu, Hawaii 96822.

The authors deeply apologize for misspelling the name of Professor Josef Dadok (not Datok) in the acknowledgment of this communication and a preceding communication on the periodate oxidation of *N*-(*p*-bromobenzoyl)palytoxin [Moore, R. E.; Woolard, F. X.; Bartolini, G. *J. Am. Chem. Soc.* **1980**, *102*, 7370].

Ortho Ester Hydrolysis: Direct Evidence for a Three-Stage Reaction Mechanism [*J. Am. Chem. Soc.* **1979**, *101*, 2669]. M. AHMAD, R. G. BERGSTROM, M. J. CASHEN, Y. CHIANG, A. J. KRESGE,* R. A. MCCLELLAND,* and M. F. POWELL, Department of Chemistry, University of Toronto, Scarborough College, West Hill, Ontario M1C 1A4, Canada.

We have detected an error in our calculation of the hydronium-ion catalytic coefficient for the hydrolysis of 2-methyl-2-methoxy-1,3-dioxolane (**11**) in biphosphate buffers:



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The correct value is $k_{H^+} = (1.50 \pm 0.03) \times 10^4$ M^{-1} s^{-1} rather than the $k_{H^+} = (1.50 \pm 0.03) \times 10^3$ M^{-1} s^{-1} originally reported. This correction requires reversal of our original conclusion that hydrolysis of this ortho ester does not undergo a change in

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